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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as P C O 2 = 0.285 torr and P N 2 = 593.525 torr. What is the partial pressure of oxygen?

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SECTION 1 continued 5. Convert a pressure of 0.200 atm to the following units: 152 a. mm Hg 20.3 b. kPa 6. When an explosive like TNT is detonated, a mixture of gases at high temperature is created. Suppose that gas X has a pressure of 50 atm, gas Y has a pressure of 20 atm, and gas Z has a pressure of 10 atm. 80 atm a. What is the total ...

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Modern Chemistry 76 Quiz Section Quiz: Gas Volumes and the Ideal Gas Law In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question. ____ 1. At the same temperature and pressure, balloons of equal volume always contain